

The solutions of extra practice in Sh 8 =



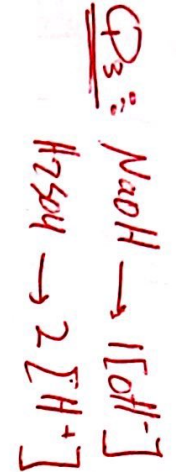
$\Rightarrow \frac{2 \text{ kmol. k}^{-1}}{\text{Base}} = \frac{1 \text{ kmol. k}^{-1}}{\text{Acid}} \Rightarrow 2 \times 0.15 \text{ k}^{-1} = 1 \times 0.29 \text{ k}^{-1}$

$V = 43.5 \text{ ml}$ ✓



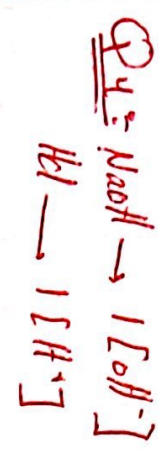
$\frac{1 \text{ kmol. k}^{-1}}{\text{Base}} = \frac{1 \text{ kmol. k}^{-1}}{\text{Acid}} \Rightarrow 83 \times 0.145 = 235 \text{ km}$

$M = 0.159 \approx 0.16 \text{ M}$ ✓



$\frac{1 \text{ kmol. k}^{-1}}{\text{Base}} = \frac{2 \text{ kmol. k}^{-1}}{\text{Acid}} \Rightarrow 38 \times 0.175 = 2 \times 155 \text{ km}$

$M = 0.10919 \approx 0.1092 \text{ M}$ ✓



$\frac{1 \text{ kmol. k}^{-1}}{\text{Base}} = \frac{1 \text{ kmol. k}^{-1}}{\text{Acid}} \Rightarrow 12.5 \times 0.3 = 285 \text{ km}$

$M = 0.0132 = \frac{\# \text{ mole}}{V = 0.285} \Rightarrow \# \text{ mole} = 0.00375 = \frac{m}{M.W. = 46}$

$M = 0.15 \text{ g}$ ✓

OR $\Rightarrow m \cdot V = \# \text{ mole} \rightarrow 12.5 \times 0.3 = 3.75 \times 10^{-3}$

$\# \text{ mole} = \frac{\text{mass}}{\text{mw}} \Rightarrow 3.75 \times 10^{-3} = \frac{m}{46} = 150 \times 10^{-5} = 0.15 \text{ g}$ ✓



$\# \text{ mole} = \frac{m \cdot k^{-1}}{1000} = \frac{99 \times 3}{1000} = \frac{99}{100} = 99 \text{ mmol} \approx 100 \text{ mmol}$ ✓

thx! b.k